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$\text{NaCl(aq)} + \text{AgNO}_3(\text{aq}) \rightarrow \text{AgCl(s)} + \text{NaNO}_3(\text{aq})$
 $\text{H}_2\text{SO}_4(\text{aq}) + 2\text{NaOH(aq)} \rightarrow \text{Na}_2\text{SO}_4(\text{aq}) + 2\text{H}_2\text{O(l)}$
 $\text{N}_2\text{O}_3(\text{aq}) + 2\text{HCl(aq)} \rightarrow 2\text{NaCl(aq)} + \text{H}_2\text{O(l)} + \text{SO}_2(\text{g})$
A compound burns in the presence of oxygen, producing energy in the form of heat and light. The combustion of organic compounds produces carbon dioxide and water.

~~Classifying Chemical Reactions—50Webs~~

Summary. In this inquiry-based experiment, the students design and carry out a sequence of chemical reactions for the separation and identification of iron, silver, and zinc ions in water. Two parallel series of tests will be carried out, one using a known sample containing all three metal cations and the other using unknown samples containing one, two, or all three cations.

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